

From

The Secretary,
Haryana State Board of Technical Education,
Bays- 7-12, Sector- 04,
Panchkula

To

The Principals of All Govt. / Aided / Pvt. Institutions offering Diploma Courses

Memo No. 3400 / Exam/HSBTE


Dated: 20.02.2020

Subject: Change in Question Paper Pattern of Semester Board Exams

In the Examination Committee Meeting, it has been decided to include Multiple Choice Questions (MCQ) henceforth in the Semester Board Exams with the condition that the total marks of objective questions / MCQ of one mark each must not exceed 20% of the total marks of the subject(s). Further 05 marks of Continuous Assessment (CA) shall be discontinued from 2020-21 academic session. Accordingly, below mentioned pattern of question papers recommended by the Examination Committee shall be followed by the Board in the Semester Board Exams:

Maximum marks of question Papers	MCQ	Objective type Questions	Short Answer Questions	Long Answer Questions	Total Questions (including options) & Total Marks
	Session A	Session B	Session C	Session D	
Pattern for 100 Marks Question Papers	Options = 10 Q 10 Q x 1 = 10 Marks	Options = 10 Q 10 Q x 1 = 10 Marks	Options = 15 Q 12 Q x 5 = 60 Marks	Options = 3 Q 2 Q x 10 = 20 Marks	Total 38 Q Total 100 Marks
Pattern for 90 Marks Question Papers	Options = 09 Q 09 Q x 1 = 09 Marks	Options = 09 Q 09 Q x 1 = 09 Marks	Options = 15 Q 12 Q x 4 = 48 Marks	Options = 4 Q 3 Q x 08 = 24 Marks	Total 37 Q Total 100 Marks
Pattern for 80 Marks Question Papers	Options = 08 Q 08 Q x 1 = 08 Marks	Options = 08 Q 08 Q x 1 = 08 Marks	Options = 10 Q 8 Q x 5 = 40 Marks	Options = 4 Q 3 Q x 8 = 24 Marks	Total 30 Q Total 80 Marks
Pattern for 60 Marks Question Papers	Options = 06 Q 06 Q x 1 = 06 Marks	Options = 06 Q 06 Q x 1 = 06 Marks	Options = 10 Q 8 Q x 4 = 32 Marks	Options = 3 Q 2 Q x 8 = 16 Marks	Total 25 Q Total 60 Marks

However, it may please be noted that the above question paper pattern shall be implemented in phased manner and therefore, in the coming Board Exams to be held in May 2020, the revised question paper pattern shall be implemented only for the 1st year (Annual Pattern) of Diploma Courses including 1st year of Pharmacy students. Sample Question Papers of 60 marks and 80 marks are attached herewith. The revised pattern for remaining subject(s) shall be implemented in subsequent Exam(s). This may please be informed to all concerned students and should also be displayed on the Notice Board / Website of the Institutions.


20/2/2020
Controller of Examination
for Secretary, HSBTE,
Panchkula

SAMPLE PAPER for May 2020 Exam

**PHARMACY (1st year)
PHARMACEUTICAL CHEMISTRY 1**

Subject Id- (2112)

M.M. 80

TIME 3 HRS.

SECTION A

Multiple choice questions. All questions are compulsory.

(08x1=8)

- Q.1 Muriatic acid is chemically___
- | | |
|-------------------|----------------------|
| a. Sulphuric acid | b. Nitric acid |
| c. Acetic acid | d. Hydrochloric acid |
- Q.2 Laughing gas is chemically___
- | | |
|------------------------|------------------------|
| a. Nitrous oxide gas | b. carbondioxide gas |
| c. carbon monoxide gas | d. sulphur dioxide gas |
- Q.3 Reducing agent used in limit test of arsenic is___
- | | |
|---------------------|----------------------|
| a. Potassium iodide | b. Granulated zinc |
| c. lead acetate | d. mercuric chloride |
- Q.4 White vitriol is___
- | | |
|---------------------|-----------------------|
| a. Copper sulphate | b. zinc sulphate |
| c. ferrous sulphate | d. magnesium sulphate |
- Q.5 Reagent used in chloride limit test is___
- | | |
|-----------------------|--------------------|
| a. Silver nitrate | b. sodium chloride |
| c. potassium chloride | d. HCl |
- Q.6 Sodium bicarbonate is used as an___
- | | |
|---------------|---------------------|
| a. antacid | b. acidifying agent |
| c. astringent | d. antioxidant |
- Q.7 Deficiency of iron leads to___
- | | |
|--------------------|-----------------|
| a. Goiter | b. anemia |
| c. Night blindness | d. hypocalcemia |

- Q.8 Sodium thiosulphate is used as a antidote in poisoning
- | | |
|------------------------------|----------------------|
| a. carbon monoxide poisoning | b. cyanide poisoning |
| c. heavy meal poisoning | d. lead poisoning |

SECTION B

Objective questions. All questions are compulsory

(8x1=8)

- Q.09 Write down the chemical formula of caustic soda.
- Q.10 Define emetics.
- Q.11 Write down the name of universal antidote.
- Q.12 Mention the use of Iron.
- Q.13 Define the term Anemia.
- Q.14 Define the term disinfectant.
- Q.15 Mention two extracellular ions used as electrolyte.
- Q.16 Write down the common name of copper sulphate.

SECTION C

Short answer questions. Attempt any 8 out of 10 questions.

(8x 5=40)

- Q.17 Define antacid. Write down the ideal characteristics of antacids.
- Q.18 Write down the storage and uses of chlorinated lime.
- Q.19 Write down the mechanism of action of astringent.
- Q.20 Write down the chemical formula and uses of following compounds:
- | | | |
|------------------|------------------|------------------|
| 1. Baking powder | 2. Muriatic acid | 3. Green vitriol |
|------------------|------------------|------------------|
- Q.21 Write a note on combination antacids.
- Q.22 Write down the principle of Heavy metal limit test.
- Q.23 Write a note on uses of radiopharmaceuticals.
- Q.24 Write down the official compounds of iron.
- Q.25 Define antidote. Write down the mechanism of action of cyanide poisoning.
- Q.26 Write a note on storage of radiopharmaceutical compound.

SECTION D

Long answer questions. Attempt any 3 Out of 4 questions. (3x8=24)

- Q.27 Write a detail note on sources of impurities in pharmaceuticals

- Q.28 Explain the limit test of arsenic with diagram, principle and procedure.
- Q.29 Define titration. Explain different types of titrations with principle.
- Q.30 Briefly describe Identification reaction for any two cations and any two anions.

SAMPLE PAPER for May 2020 Exam

1st Year (Annual Pattern)
Subject: Applied Chemistry

Subject Id: 180014

Time Allowed: 3Hrs**MM: 60**

Section-A

Multiple choice questions. All are compulsory.

(06x1= 06)

- Q.1 Substance in its molten or aqueous solution states conduct electricity is called _____
a) Metal
b) Electrolyte
c) Semiconductor
d) None of these
- Q.2 The ore of Copper is
a) Haematite ($\text{Fe}_2\text{O}_3 \cdot x\text{H}_2\text{O}$)
b) Bauxite ($\text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$)
c) Copper Pyrite (CuFeS_2)
d) Ferro chrome (FeOCr_2O_3)
- Q.3 The water which easily dissolves soap in it is known as _____
a) Sea water
b) Hard water
c) Soft water
d) None of these
- Q.4 The substance which is added between sliding surface to decrease friction and prevent them from wear and tear is called _____
a) Acid
b) Base
c) Salt
d) Lubricant
- Q.5 The unit of Calorific value of fuel is
a) Neuton
b) Kelvin
c) Decibel
d) Cal/g
- Q.6 The Acidity of good lubricant should be
a) High
b) Low
c) Moderate
d) None of these

Section-B

Objective questions. All are compulsory.

(06x1= 06)

- Q.7 The fire point of lubricant is always _____ than flash point (Higher/ Lower).
- Q.8 The process of decomposition of electrolyte on passing electricity from its aqueous or molten state is called electrolysis. (True/ False)
- Q.9 The hardness of water which can be easily removed by boiling water is called _____.
- Q.10 The electrolyte which dissociate almost completely into its ions when dissolve in water or melted is called _____.
- Q.11 Give any two example of metals.
- Q.12 Give two examples of liquid fuel.

Section-C

Short answer questions. Attempt any eight questions out of ten questions.

(08x4= 32)

- Q.13 Write down any four purposes of alloying?
- Q.14 What do you mean by “electro refining”? Explain with the help of suitable diagram.

- Q.15 Classify fuel on the basis of their physical state.
- Q.16 Define i) Metallurgy ii) Fuel.
- Q.17 What do you mean by hard water and soft water?
- Q.18 Define the terms flash point and fire point.
- Q.19 Enlist four advantages of gaseous fuel over solid fuels?
- Q.20 Differentiate (two) between roasting and calcination.
- Q.21 Explain the froth floatation method for concentration of ore.
- Q.22 Enlist four disadvantages of caustic embrittlement.

Section-D

Long answer questions. Attempt any two questions out of three questions.

(2x8= 16)

- Q.23 a) Write down any four difference between thermoplastic and thermosetting polymer?
b) Write down any four uses of polymer
- Q.24 a) Write down any for characteristics of potable (drinking) water.
b) What are the causes of hardness of temporary and permanent hardness of water?
- Q.25 a) Write down any for function of lubricant.
b) Explain the term total acid value and emulsification.